

# Thermochemistry Practice Calculations Worksheet Answers

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## Thermochemistry Practice Calculations Worksheet Answers

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Thermochemistry Worksheet #1 1. The reaction of magnesium with sulfuric acid was carried out in a calorimeter. This reaction caused the temperature of 27.0 grams of liquid water, within the calorimeter, to raise from 25.0 C to 76.0 C. Calculate the energy associated with this reaction. 2.

### **Thermochemistry Worksheet #1**

Acces PDF Thermochemistry Worksheet 2 Answers by: USER Created Date: 5/23/2007 1:03:00 PM Company: River Dell Regional Schools Thermochemistry Problems - Worksheet Number One Answers, Thermochemistry Practice Problems 2 1 6. When 26.7 g of H<sub>2</sub>S was burned in excess oxygen, 406 kJ was released.

### **Thermochemistry Worksheet 2 Answers**

THERMOCHEMISTRY CALCULATIONS WORKSHEET 1 Using reaction equation ratios 1. How much heat will be released when 6.44 g of Sulfur reacts with excess O<sub>2</sub> according to the following equation?  $2S + 3O_2 \rightarrow 2SO_3 \Delta H^\circ = -791.4\text{kJ}$  2. How much heat will be released when 4.72 g of Carbon reacts with excess O<sub>2</sub> according to the following equation?  $C + O_2 \rightarrow CO_2$

### **THERMOCHEMISTRY CALCULATIONS WORKSHEET 1**

Show all the work for the calculations and give the answers in the correct significant figures. 1) Indicate if the following reactions are endo or exothermic. CH<sub>4</sub>

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S. Thermochemistry Practice Problems ANSWER KEY 1. Calculate the standard enthalpy of reaction for the

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## **Thermochemistry With Answers - Teacher Worksheets**

Answers, Thermochemistry Practice Problems 2 1 6. When ... This problem is essentially identical to the calculations you did in lab with your data from reacting HCl with NaOH (Expt. 14, part C)!! As in the lab, to estimate an answer to this problem using the information given, one must assume that the density of the solutions is 1.00 g./mL ...

## **Answers, Thermochemistry Practice Problems 2**

Thermochemistry Exam1 and Problem Solutions 1. Which ones of the following reactions are endothermic in other words  $\Delta H$  is positive? I.  $\text{H}_2\text{O}(\text{l}) + 10,5\text{kcal} \rightarrow \text{H}_2\text{O}(\text{g}) \Delta H1$  II.  $2\text{NH}_3 + 22\text{kcal}$

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## Thermochemistry Worksheet 1 Answers

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## Thermochemistry Worksheet 1 Answers - PvdA

This is the same answer you obtained in Q1, as expected on the basis of Hess's Law. But notice that the answer we obtain here is the following sum, where  $\Delta H^{\circ}_f(\text{H}_2) = 0$ , because  $\text{H}_2(\text{g})$  is an element in its standard state:

$$\begin{aligned} \Delta H^{\circ}_{\text{rxn}} &= \Delta H^{\circ}_f(\text{C}_2\text{H}_4) - \left[ \Delta H^{\circ}_f(\text{C}_2\text{H}_2) + \Delta H^{\circ}_f(\text{H}_2) \right] \\ &= +52.3 \text{ kJ} - \left[ -226.8 \text{ kJ} + 0 \text{ kJ} \right] \end{aligned}$$

...

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