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At 450 °C, the value of the equilibrium constant for the following system is 6.59×10^{-3} . If $[\text{NH}_3] = 1.23 \times 10^{-4}\text{M}$ and $[\text{H}_2] = 2.75 \times 10^{-2}\text{M}$ at equilibrium,

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determine the concentration of N_2 at that point. $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$

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The chemical equilibrium can only occur at temperatures above room temperature. The state of equilibrium can be maintained over time as long as

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all factors change. The rate of the forward reaction...

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a reversible reaction in which there is no longer any change in the concentration of reactants and products. chemical equilibrium. a state of balance in which

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the rate of a forward reaction equals the rate of the reverse reaction and the concentrations of products and reactants remain unchanged.
equilibrium.

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14.4 moles of A are mixed with 4 moles

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of B. At equilibrium for the reaction $A + B \rightleftharpoons C + D$, 2 moles of C and D are formed. The equilibrium constant for the reaction will be: a) $\frac{1}{4}$ b) $\frac{1}{2}$ c) 1. d) 4. ANS:

Chemical Equilibrium Important Questions And Answers

Chapter 18 Chemical Equilibrium

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Chapter 18 Chemical Equilibrium Worksheet Answers

SHORT ANSWER Answer the following questions in the space provided. 1. Write the equilibrium expression for the following hypothetical equation: $3A(aq) + B(aq) \rightarrow 2C(aq) + 3D(aq)$ 2.

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Chemical Equilibria and Reaction
Quotients Many chemical reactions don't
just go one way, they go forwards and
backwards. Once there is balance
between the two, this is Chapter 15 -

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Chemical Equilibrium: Part 3 of 12 In this continued lecture on chemical equilibrium, I'll teach you

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Forward: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ Backward:
 $\text{N}_2 + 3\text{H}_2 \leftarrow 2\text{NH}_3$ The equation is typically combined like so: $\text{N}_2 + 3\text{H}_2$

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----> 2NH_3 Equilibrium means that opposing processes are in balance. Reversible reactions balance each other because they take place at equal rates. This is called chemical equilibrium.

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equilibrium to shift towards the reactants to create more $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$.

- Adding 3mL of 12M HCl puts stress on the reactants side (adding Cl^-), which forces to equilibrium to shift towards products/right to create more $[\text{CoCl}_4]^{2-}$.

- Adding deionized water puts a stress on H_2O on the product side and ultimately decreases

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Lab 5- Chemical Equilibrium and Le Chatelier's Principle ...

Equilibrium system and stress: State Le Chatelier's principle. Apply Le Chatelier's principle for stress. Discuss the common-ion effect. Discuss the practicalness of Le Chatelier's principle. Holt Do now Page 512: Define Le Chatelier's

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principle. Page 518: Answer questions 1-13. Page 521-522: Answer questions 17, 18, 19

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