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1km 10 m 10 m 10 m m 10 m. = = 33 6 9() (μ) μ The given measurement is 1.0 km (two significant figures), which implies our result should be written as $1.0 \times 10^9 \mu\text{m}$.

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E1-22 First find the "logarithmic average" by $\log d_{av} = \frac{1}{2} \log(2 \cdot 10^{26}) + \log(1 \cdot 10^{15})$ $\frac{1}{2} \log 2 \cdot 10^{26}$
 $\frac{1}{2} \log 2 \cdot 10^{11} = \log p \cdot 2 \cdot 10^{11}$ Solve, and $d_{av} = 450 \text{ km}$. E1-23 The number of atoms is

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given by $(1 \text{ kg}) / (1.00783 + 1.661 \times 10^{-27} \text{ kg})$, or 5.974×10^{26} atoms.

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