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Chapter 21 Notes Electrochemical Cells Page 1 of 4 21.1 Electrochemical Cells Electrochemical Processes o A Spontaneous Redox Reaction KEY = For any two metals in an activity series, the more active metal is more readily oxidized than the less active metal o Redox Reactions and Electrochemistry

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1. In an electrolytic cell the electrode at which the electrons enter the solution is called the \_\_\_\_ ; the chemical change that occurs at this electrode is called \_\_\_\_\_. (a) anode, oxidation (b) anode, reduction (c) cathode, oxidation (d) cathode, reduction (e) cannot tell unless we know the species being oxidized and reduced.

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### **Chapter 21 Electrochemistry Answers**

Chapter 21: ELECTROCHEMISTRY TYING IT ALL TOGETHER On the basis of these marks the correct answer will be. (a) In dilute sulphuric acid solution, hydrogen will be reduced at ... Electrochemistry - Answers Answer: (a) tin electrode is the cathode; cathode is the site of reduction

### **Electrochemistry Problems And Answers**

21-1 Chapter 21 Electrochemistry Section 21.1. Electrochemistry represents the interconversion of chemical energy and electrical energy. Electrochemistry involves redox (reduction-oxidation) reactions because the electrical energy (flow of electrons) has at its origin the oxidation (loss of electrons) and reduction (gain of electrons) of species.

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